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$2\text{N}_2\text{O}(g) \rightarrow 4\text{NO}(g) + \text{O}_2(g)$

Time (s)	Moles N_2O	Moles NO	Moles O_2
0	0.4720	0	0
600	0.4194	0.0252	0.0063
1200	0.4104	0.0432	0.0108
1800	0.4032	0.0576	0.0144

Handwritten notes: $2(0.0626)$, $4(0.0036)$, 0.009 , $+ 0.018$, $+ 0.0045$

Critical Thinking Questions

7. Calculate the change in moles of each reactant and product between 600 and 1200 seconds.

change in moles of N_2O : $2:4$ change in moles of NO : $4:1$ change in moles of O_2 :
 $= 0.009 + 0.018 + 0.0045$

8. Compare your answers in question 7. What is the relationship between the coefficients in the balanced chemical equation and the number of moles used up or produced in a reaction?

9. Fill in the missing blanks in the table above.

10. Calculate the following values. Be sure to use molarity in your calculations.

a) Rate of disappearance of N_2O between time 0 and 600 s.

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